Empirical formula Example:-

Tuesday, November 11, 2025 9:54 AM

Question:- A compound has 24g of Carbon, 8g of Hydrogen, 32g of oxygen.

Solution:-

Divide each mass by it Atomic mass.

$$\text{Carbon} = \frac{24}{12} = 2.$$

$$\text{Hydrogen} = \frac{8}{1} = 8$$

$$\text{Oxygen} = \frac{32}{16} = 2$$

Find the simplest Ration.

C : H : O

2 : 8 : 2

1 : 4 : 1

CH_4O ⟶ Answer. (Empirical formula).