Chemistry

Covalent bonding

Covalent bond is formed by sharing of electrons between two non metals

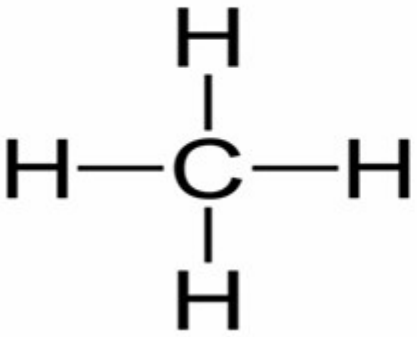
Atom share electron to get a whole outer shell and becomes stable like noble gases

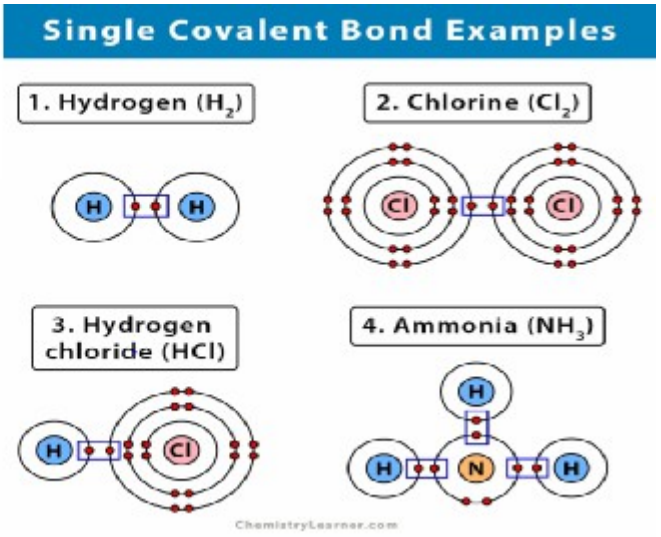
Types of covalent bonding

1. Single covalent bond
2. double covalent bond
3. triple covalent bond

Single covalent bond

- It has a single bond pair
- 1 pair of share electrons
- it is represented by one dash (-)

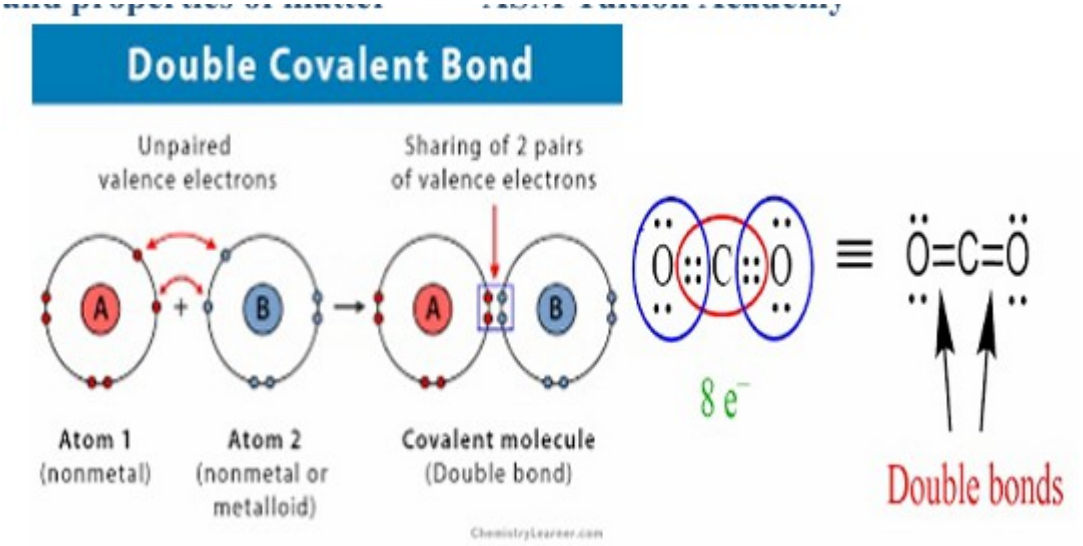




Double covalent bond

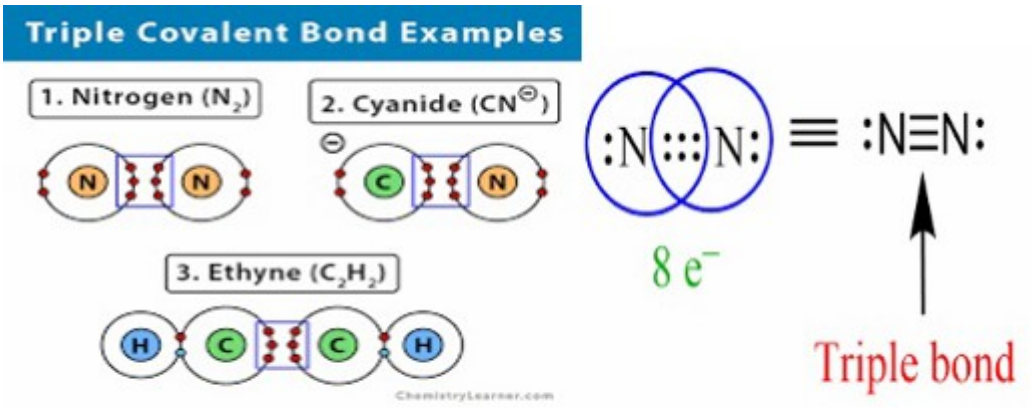
- Double covalent bond means two pair share electrodes

- it is represented by 2 dashes
- double covalent bonds are much stronger than single bonds but they are less table



Triple covalent bond

- triple covalent bond means 3 share pair of electrons
- triple bonds are represented by three dash is
- they are least stable they type of covalent bond



➤ **Properties of simple molecular substances**

Properties of simple molecular substances

- Low melting and boiling point, because they have no free electrons and no overall electric charge in any state
- Intermolecular forces, hydrogen Ammonia Methane and pure water are also simple molecules
- They all have strong covalent bonds inside the molecules but much weaker intermolecular force between the molecules
- At room temperature simple molecular substances are gases solids liquids with low melting and boiling point
- Physical properties are that they are soft because the intermolecular forces are weak and they are easy to break